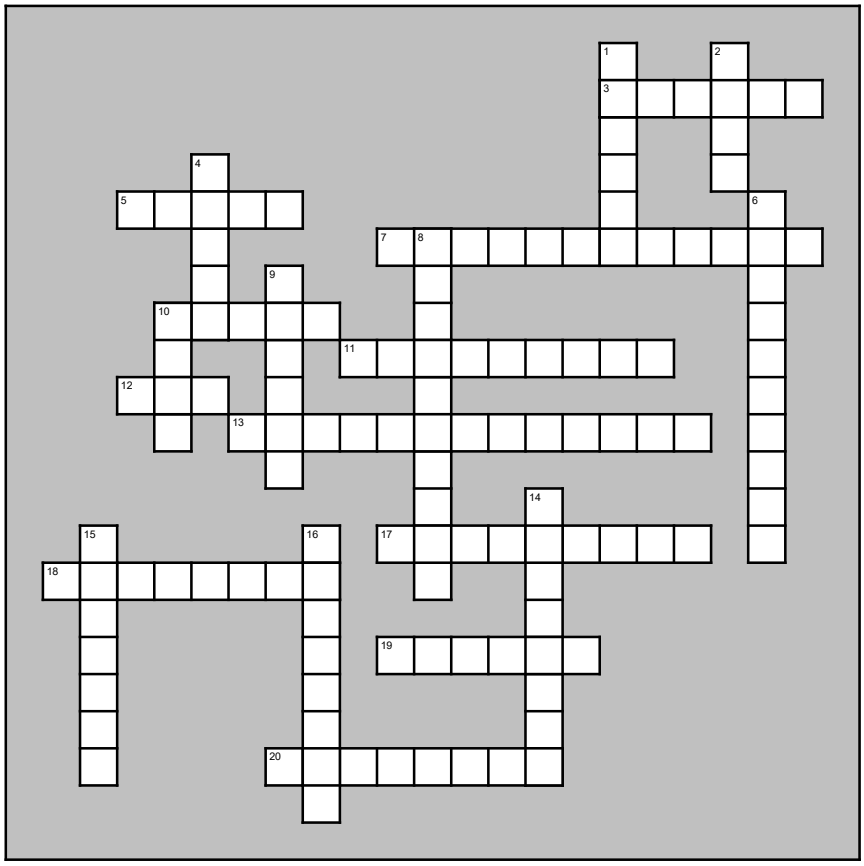
Crossword



Across

3. Metals that have 1 valence electron, forming positive ions and bonding with non-metals. Very strong reactivity. Lithium, Sodium.
5. The vertical columns on the Periodic Table of elements in which elements have similar properties; also called "families."
7. The modern Periodic Table is arranged according to ___ ___. This indicates the number of protons in an atom.
10. The number of periods on the Periodic Table of Elements.
11. Arranged the elements on the Periodic Table into rows in order of increasing mass so that elements with similar properties were in the same column.
12. Atomic mass unit: signifying the mass of one atom of an element. Based on the carbon 12 atom.
13. Metals that have 2 valence electrons, forming positive ions and bonding with non-metals. Magnesium, Calcium.
17. Gain electrons easily and therefore form negative ions. Found on the right side of the Periodic Table. Brittle, non-conductors.
18. Non-metals that have 7 valence electrons, forming negative ions and bonding with metals. Very strong reactivity. Fluorine, Chlorine.
19. The horizontal row on the Periodic Table of Elements in which elements vary by increasing atomic number.
20. The elements on the LEFT side of the Periodic Table, these properties become stronger. Most elements are this.

Down

1. All living things contain this element. Most of our body is composed of it. It has an atomic number of 6 and an atomic mass of 12. Basis for the amu.
2. Atomic ___ was the former way to arrange the elements on the Periodic Table. An important factor in studying Chemical reactions. Measured in amu's.
4. Gases in the last group of the Periodic Table that are extremely unreactive elements; colorless, odorless.
6. "Staircase elements" on the Periodic Table that possess both metallic and non-metallic properties. Silicon, Arsenic.
8. Elements, usually classified as metals, in the middle of the Periodic Table whose properties can vary because their valence varies; distinctive colors. e.g. Copper, Mercury, Silver.
9. Lose electrons easily and therefore form positive ions. Found on the left side of the Periodic Table. Malleable, ductile, shiny, hard, good conductors.
10. Elements in a group possess the ___ number of valence electrons, causing similar properties for all elements in that "family."
14. A law in which the properties of elements changing consistently along a row of the Periodic Table (based on atomic numbre) and then repeats on the next row.
15. The outermost electrons in any atom ... usually the electrons involved in bonding. Elements are grouped because they have the same number of these.
16. Atoms of the same element, having a different number of neutrons. e.g Cl-35 has 17 protons, 18 neutrons ... Cl-37 has 17 protons, 20 neutrons.