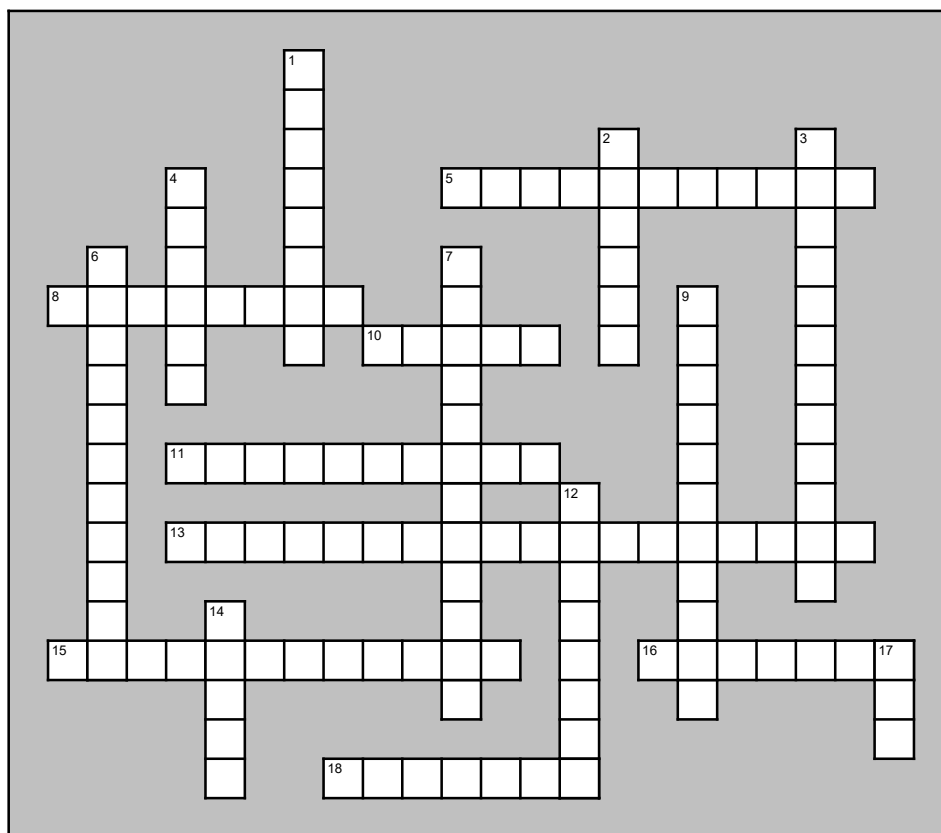


# Crossword



## Across

5. Phase change from solid to gas without becoming a liquid. Energy is absorbed (endothermic). Physical change. e.g. dry ice forms clouds. Reverse of deposition.
8. Describe attributes or characteristics of substances or mixtures which do NOT change that substance or mixture. Density, texture, phase changes.
10. Pressure caused by collisions of vapor molecules above a liquid. The tendency to boil. As temperature increases, \_\_\_ pressure increases, meaning it is easier to boil at a higher temperature.
11. Phase change from gas to solid without becoming a liquid. Energy is released (exothermic). Physical change. Reverse of sublimation.
13. The amount of energy gained or lost at a substance's boiling/condensation point to phase change a liquid into gas (endothermic) or a gas into a liquid (exothermic). Varies from substance to substance.
15. The amount of energy gained or lost at a substance's melting/freezing point to phase change a solid into liquid (endothermic) or a liquid into a solid (exothermic). Varies from substance to substance.
16. Endothermic phase change from liquid to gas. The vapor pressure of the liquid is equal or greater than the atmospheric pressure. The lower the atmospheric pressure, the greater the tendency. Reverse of condensation.
18. Phase change from solid to liquid. Energy is absorbed (endothermic). Physical change when molecules gain enough energy to overcome attractions and move from their fixed positions. Reverse of freezing.

## Down

1. Properties producing a change in the composition of matter into a new substance. e.g flammability, reactivity. Commonly have a color change, form a gas or a precipitate.
2. The phase with definite shape and indefinite volume. It has a moderate molecular motion. Takes the shape of its container.
3. Phase change from liquid to solid. Energy is released (exothermic). Physical change. e.g. dew. Reverse of boiling.
4. 99% of all matter in the universe is thought to be this state of matter.
6. The reversible physical change that occurs when a substance changes from one state of matter to another.
7. Phase change from liquid to gas. Energy is absorbed (endothermic). Physical change when molecules gain enough energy to overcome attractions and move most randomly. e.g. evaporation and boiling.
9. Liquid becomes gas. Takes place at the surface of a liquid and can occur at temperatures below the boiling point.
12. Phase change from liquid to solid. Energy is released (exothermic). Physical change when molecules lose enough energy to become attracted and drawn into a more orderly arrangement. Reverse of melting.
14. The phase with definite shape and volume and the lowest molecular motion.
17. The phase with NO definite shape or volume and the highest molecular motion. Fills its container.